

quarter century. A very large fraction of this growth has been in the area of inorganic solids, of which metal oxides are the most important. This growth in metal oxides has been soundly based on theory, supported by immediate technological application, and has resulted in the proliferation of new compounds many of which have been, literally, made to order. It might have been thought that the ingenious scientists responsible for this development would have been people who refer to themselves as inorganic chemists—but not so. Of the approximately 1000

names in the author index only a mere handful are known to this reviewer as inorganic chemists, and three of these with impeccable credentials (M. Calvin, G. N. Lewis, and T. Moeller) actually appear in disguise as Colvin (p. 519), S. G. N. Lewis (p. 544), and Moeller (p. 2).

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### Erratum

Vol. 10, No. 4 (1968), in the article, "Studies on Pore Systems in Catalysts. XIV. Calculation of the Cumulative Distribution Functions for Slit-Shaped Pores from the Desorption Branch of a Nitrogen Sorption Isotherm," by J. C. P. Broekhoff and J. H. de Boer, pp. 391–400:

Page 395, Eq. (15b) should read:

$$\frac{d}{2} - t_a = \frac{2.02^5 + 16.11(1/t_a - 2d) + 1.483[\exp(-0.05685d) - \exp(-0.1137t_a)]}{\log(p_o/p_D)} \quad (15b)$$